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Research Article

PREPARATION AND *IN VITRO* EVALUATION OF IBUPROFEN MICROSPHERES USING IONIC GELATION METHOD

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ABSTRACT

Introduction: Microspheres can be defined as a matrix system that allows the drug to be homogenously dispersed, dissolved or suspended. **Aim:** Development and *in vitro* evaluation of ibuprofen floating microspheres. **Methods**: The microspheres were formulated using inotropic gelation method. A homogenous polymer solution was prepared by dissolving sodium alginate (1 g) and the polymers (1 g each) in 32 ml of distilled water. The drug was added to the polymer solution and stirred continuously to form a viscous dispersion. A 10 % w/v $Cacl_2$ solution was prepared and used as a cross–linking agent. The dispersion was added drop wise for 15 minutes for the curing reaction to take place. The spheres obtained were then washed and dried at 45 ^oC for 12 hours. **Results**: The percentage yield of the formulations ranged from 72.00 % ± 1.41 to 85.50 % ± 3.54. The angle of repose ranged from 3.49^o ± 0.04 to 9.49^o ± 0.31, while the bulk density ranged from 0.60 g/ml ± 0.00 to 0.41 g/ml ± 0.00. The Carr's index ranged from 13.04 % ± 0.04 to 14.50 % ± 0.35. The swelling index ranged from 65.50 % ± 0.35 to 85.00 % ± 0.71. The *in vitro* drug release showed that formulation F-5 gave the least release at 28.80 % ± 0.85 after 4 hours without a significant difference ($p < 0.05$). **Conclusion:** The idea of formulating floating microspheres containing ibuprofen gave a suitable practical approach that achieved a prolonged therapeutic effect by releasing the active drug over an extended period of time.

KEYWORDS: Ibuprofen, Ionic gelation, Microspheres, Gastrointestinal infection

1. INTRODUCTION

Microencapsulation is the process by which solids, liquids and gases are enclosed in microscopic particles by formation of wall coatings around the drug.[1] Microspheres are small spherical particles within the 1-100 µm range.[2] Microspheres can be characterized as a matrix system that allows the drug to be homogenously dispersed, dissolved or suspended.[3] There are different techniques involved in the production of microspheres. The solvent evaporation method is used were the polymer is dissolved in an organic solvent and the drug is either dissolved or dispersed in a polymer solution. The solution containing the drug is then emulsified into an aqueous phase containing suitable additive to form oil in water emulsion.[4] The ionotropic gelation method is based on the ability of the polyelectrolyte to cross link in the presence of counter ions in order to form beads.[5] The emulsion solvent diffusion method is the process were the drug is dissolved in the organic solvent and the solution is dispersed in the aqueous solvent producing the emulsion droplets.[6] In single emulsion technique, natural polymers are dissolved or dispersed in aqueous medium using a cross-linking agent.[7] Gastroretentive floating microspheres are known low density systems that have the ability to maintain buoyancy to float over a prolonged period of time.[8] There are various types of microspheres that are used for drug delivery. They include the bioadhesive microspheres, magnetic microspheres, polymer microspheres, radioactive and floating microspheres.[9] Advantages associated with floating microspheres include: enhanced bioavailability, enhanced biotransformation, sustained drug delivery, minimized adverse reactions and site specific drug delivery.[10] An ideal microsphere is meant to possess some unique

characteristics such as non-toxicity, relative stability, increase therapeutic efficiency and biocompatibility.[11] Ibuprofen which is a commonly used non-steroidal anti-inflammatory drug (NSAID) has some major challenges especially in oral administration. Its absorption occurs in the upper gastrointestinal tract, leading to gastric irritation and ulceration. Although it has short half-life which necessitates frequent dosing, affects patience compliance and increases adverse effect. This study was based on the preparation and *in vitro* evaluation of ibuprofen floating microspheres using different polymers in order to achieve controlled drug delivery and a focus on optimizing drug release kinetics and enhancing therapeutic efficacy.

2. MATERIALS AND METHODS

Ibuprofen (IBU/2/2004/0072A), was purchased from Emzor Pharmaceuticals Ltd, Lagos. Sodium alginate, ethyl cellulose, sodium carboxy methyl cellulose (Na-CMC), were purchased from (Sigma Aldrich, Kosher, USA). Methanol was obtained from (Astron Chemicals, Ahmedabad). Glycerin and sodium hydroxide were provided by (Mingtai Chemical Taiwan). Calcium chloride was obtained from (Evonik, Germany). Hydroxy propyl methyl cellulose (HPMC) were obtained from (DFE Pharma, UK), Sorbitol was obtained from (TCI, USA). Distilled water was obtained from (UNN Water Resources Management Laboratories Ltd; UNN, Enugu State, Nigeria). All chemicals used were of analytical grade.

2.1 Method of preparation

Orifice inotropic gelation method was used for the preparation of ibuprofen microspheres using polymers such as ethyl cellulose (EC), sodium carboxyl methyl cellulose (Na-CMC), hydroxyl propyl methyl cellulose (HPMC) and sodium alginate.[12] A homogenous polymer solution was prepared by dissolving sodium alginate (1g) and the polymers (1g) in purified water (32 ml). Ibuprofen (1g), the active substance was added to the polymer solution and stirred thoroughly to form a viscous dispersion. A 10 % w/v quantity of calcium chloride solution was prepared which was used as a cross linking agent. The prepared dispersion was then manually added drop wise into calcium chloride (10 % w/v) solution (40 ml) using a syringe having a needle of size (no.18). The calcium chloride solution having the droplets was then allowed to stay for 15 minutes for the curing reaction to take place and produce spherical rigid drug loaded spheres. The spheres obtained after the reaction were then collected and washed repeatedly with acetone. After washing, the spheres were properly dried at 45^{o}C for 12 hours.

Table 1: Composition of formulations.

Ingredients (g)/Batches	F-1	F-2	F-3	F-4	F-5	F-6	F-7	F-8	F-9
Ibuprofen	1	1	1	1	1	1	1	1	1
Sodium alginate	1	1	0.8	1	0.8	1	0.8	1	0.8
Ethyl cellulose	1	0.2	1	-	-	-	-	-	-
Na-CMC	-	-	-	1	0.2	1	-	-	-
HPMC	-	-	-	-	-	-	1	0.2	1
Acetone (ml)	30	30	30	30	30	30	30	30	30
$Cacl_2$	10	10	10	10	10	10	10	10	10

2.2 Yield analysis of the recovered microspheres

The relative yield was calculated based on the amount of microspheres of each formulation obtained relative to the amount of solid materials used in the dispersed phase.[13] The percentage yield was calculated according to the following equation:

$$\text{Yield (\%)} = \frac{\text{Actual weight of microspheres}}{\text{total weight of drug and polymer}} \times 100 \quad \ldots\ldots 1$$

2.3 Pre-compression evaluation of powder blend

2.3.1 Angle of repose

A sheet of fibre board was placed below the funnel orifice making sure it fits tightly. A given quantity of the microsphere (30 g) was transferred into the funnel. The fibre sheet was drawn away and the timer simultaneously started. The timer was stopped when all of the powder had passed through the funnel. The height of the heap was measured using a graduated ruler. A pencil was used to outline the base of the contour. The angle of the conical heap so formed was determined from equation 2. The powder was returned to the funnel and the experiment was repeated thrice.[13]

$$\text{Tan } \theta = \frac{\text{height of powder heap, (h)}}{\text{radius of powder heap, (r)}} \quad \ldots\ldots\ldots\ldots 2$$

2.3.2 Bulk density

This is the ratio between a given mass of powder and its bulk volume. A weighed quantity of the microsphere (30.0 g) was placed in a 100-ml graduated cylinder. The cylinder was gently dropped onto a wooden surface three times from a height of one inch at 2 sec interval. The volume assumed after the treatment was taken as the bulk volume. The experiment was repeated.[13]

$$\text{Bulk density (g/ml)} = \frac{mass}{bulk\ volume} \quad \ldots\ldots\ldots\ldots 3$$

2.3.3 Tapped density

This is the ratio between a given mass of powder and its bulk volume. A weighed quantity (30.0 g) of the powder was placed in a 100-ml graduated cylinder. The cylinder was tapped up to 500 times on the wooden surface or to a

constant volume. The final volume attained represents the tapped volume. The experiment was repeated thrice.[13]

Tapped density (g/ml) = $\frac{mass}{tapped\ volume}$ 4

2.3.4 Carr's index

This is used to access the flowability of a powder. The Carr's compressibility index (CI %) was calculated from the poured (bulk density) and tapped densities. CI was calculated using the following equation:

Carr's index = $\frac{Tapped\ density - bulk\ density}{Tapped\ density}$ x 100 5

2.3.5 Hausner's ratio

The Hausner's ratio (HR) is the ratio of tapped to bulk densities. It is a common technique widely used to describe the packing behavior of powders when they are subjected to tapping [13]

Hausner's ratio = $\frac{\text{tapped density}}{\text{bulk density}}$ 6

2.3.6 Swelling index

The weight of the microspheres was taken and then dispersed in phosphate buffer (pH 7.2) for 12 hours. The excess liquid was removed using blotting paper and the weight of the swollen microspheres taken. The swelling index was calculated thus.[14]

Swelling index = $\frac{\text{Weight of swollen microspheres} - \text{weight of dried microspheres}}{\text{weight of swollen microspheres}}$ 7

2.3.7 Drug content

A 1g quantity of sample was taken and dissolved in 100 ml distilled water in a beaker. After 24 hours, the sample was filtered and suitable dilution was done. Then the absorbance of the solution was measured at 215 nm and drug content was calculated.[15]

2.3.8 *In vitro* analysis

A 500 ml quantity of distilled water was placed in the dissolution apparatus (USP apparatus type-II paddle method). The sample was then placed in the vessel and the apparatus was operated for 4 hours at 50 resolution per minute (RPM). At a definite time interval, 5 ml was withdrawn from the vessel and another 5 ml of the blank was added to the vessel. The withdrawn fluid was then filtered and suitable dilutions were done. Samples were analyzed under UV Spectrophotometer at 277 nm.[16]

2.3.9 Morphology of the ibuprofen microspheres

The morphology of the obtained microspheres was examined under a light microscope (Zeiss, Me 63 C, West Germany) with varied magnification powers. One drop of the freshly prepared microsphere suspension was poured onto a slide and sealed with a cover glass. Photomicrographs were captured using Samsung digital camera.[14] The morphology, size, uniformity and aggregation or coalescence of the microspheres were studied.[17]

2.3.10 Drug-excipient compatibility study (FTIR spectroscopy)

Infra-red spectra of pure drug, carrier and coating materials were obtained by (Shimadzu 8400S Japan) FT-IR spectrometer. The samples were previously ground and mixed thoroughly with potassium bromide (KBr), an infra-red transparent matrix at 1:5 (sample: KRr) ratio respectively. The KBr discs were prepared by compressing the powders at a pressure of 5 tons for 5 minutes in a hydraulic press. The scans were obtained at a resolution of 4 cm^{-1} from 4000 to 400 cm^{-1}.[18]

2.4 Data analysis

All the measurements were repeated at least thrice and the data obtained analyzed by Student *t*-test and One-Way Analysis of Variance (ANOVA). Statistical analysis was performed using Statistical Product and Services Solution software (SPSS, version 22.0 Inc., Chicago IL, USA) and Excel Microsoft Office version 2012. The results were presented as mean ± SD, and statistical differences between means considered significant at ($p < 0.05$).

3.0 RESULTS AND DISCUSSION

3.1 Percentage yield of the ibuprofen microspheres

The percentage yield of the ibuprofen microspheres varied from 72.00 % ± 1.41 to 85.50 % ± 3.54 at different concentrations of the drug-polymer ratios. Batches F-7 recorded the lowest percentage yield at 72.00 % ± 1.41 without any significant difference ($p < 0.05$). According to Trivedi et al, the reduction in the percentage yield with increasing drug-polymer ratio may be due to the loss of smallest particles during filtration and washing.[15]

Table 2: Percentage yield of ibuprofen microspheres.

Formulation code	Yield (%) ± SD
F-1	85.5 ± 2.47
F-2	85.5 ± 3.54
F-3	73.0 ± 2.12
F-4	79.5 ± 5.30
F-5	74.5 ± 0.35
F-6	84.5 ± 1.06
F-7	72.0 ± 1.41
F-8	83.5 ± 1.06
F-9	82.5 ± 1.77

3.2 Angle of repose

The angle of repose is an indicator of the internal friction or cohesion between particles.[17] The results showed the angle of repose for the ibuprofen floating microsphere formulations ranged from 3.49° ± 0.04 for batch F-5 to 9.49° ± 0.31 for batch F-3 without a significant difference ($p < 0.05$). These low values indicated good flow and cohesive properties of the powders.[17]

3.3 Bulk density

Bulk density provides a measure of the flow properties of a powder, which is influenced by particle size and distribution.[17] A higher bulk density correlates with better flow characteristics. The bulk density values obtained were in the range of 0.60 g/cm^3 ± 0.00 for batch F-3 to 0.41 g/cm^3 ±0.00 for batch F-2 without a significant difference ($p < 0.05$).

3.4 Tapped density

Tapped density is dependent on particle size and size distribution. The tapped density values ranged from 0.56 g/cm^3 ± 0.01 for batch F-3 to 0.65 g/cm^3 ± 0.02 for batch F-9 without a significant difference ($p < 0.05$).

3.5 Carr's index

The Carr's index (CI) indicates the flow properties of a powder. Values between 5-15 % represents excellent flow and 12-16 % represents good flow according to BP specifications [17]. Powders with CI above 38 % are considered very poor flowing and cohesive. The CI values for the microsphere formulations were between 13.04 % ± 0.04 (batch F-6) and 14.50 % ± 0.35 (batch F-5), without a significant difference ($p < 0.05$), suggesting good to excellent flow properties.

3.6 Hausner's ratio

The Hausner's ratio (HR) provides another measure of powder flow, with a ratio between 1.00-1.11, indicating excellent flow and 1.12-1.18, representing good flow properties. HR values above 1.6 are characteristic of very poor, cohesive flow powders/granules. The HR ranged from 1.03 ± 0.02 for batch F-4, showing excellent flow, to 1.36 ± 0.06 for batch F-9, without any significant difference ($p < 0.05$), still within the range for good flow according to the British Pharmacopeia specifications.[17]

Table 3: Summary of the angle of repose, bulk density, tapped density, Carr's index, and Hausner's ratio.

Formulation code	Angle of repose (°)	Bulk density (g/ml)	Tapped density(g/ml)	Carr's Index (%)	Hausners ratio
F-1	7.19 ± 0.08	0.54 ± 0.02	0.56 ± 0.01	14.4 ± 0.26	1.29 ± 0.35
F-2	3.72 ± 0.06	0.50 ± 0.00	0.62 ± 0.01	13.6 ± 0.35	1.05 ± 0.02
F-3	9.49 ± 0.31	0.51 ± 0.01	0.60 ± 0.00	14.3 ± 0.13	1.09 ± 0.11
F-4	4.85 ± 0.02	0.41 ± 0.00	0.51 ± 0.01	15.0 ± 0.03	1.03 ± 0.02
F-5	3.49 ± 0.04	0.50 ± 0.00	0.60 ± 0.00	14.5 ± 0.35	1.17 ± 0.10
F-6	4.55 ± 0.04	0.60 ± 0.00	0.64 ± 0.01	13.04 ± 0.04	1.15 ± 0.01
F-7	5.05 ± 0.02	0.61 ± 0.00	0.63 ± 0.01	13.39 ± 0.22	1.13 ± 0.01
F-8	6.76 ± 0.06	0.53 ± 0.01	0.64 ± 0.03	13.08 ± 0.29	1.27 ± 0.13
F-9	6.87 ± 0.02	0.63 ± 0.02	0.65 ± 0.02	13.8 ± 0.11	1.36 ± 0.06

3.7 Swelling index and Drug content of ibuprofen microspheres

The ibuprofen microspheres floated for prolonged period of time when it was immersed on the dissolution medium. The percentage of the swelling index was highest for formulations with HPMC and sodium alginate polymers. There could be a direct relationship between the increase in polymer concentration and increased buoyancy time.[18] According to Chintapalli *et al*, an increase in polymer concentration led to an increase in the buoyancy time.[18] The swelling index was used to determine the amount of phosphate buffer absorbed by the microspheres after dissolving them in the buffer. From the results obtained, it ranged from 65.50 ± 0.35 to 85.00 ± 0.71 for batches F-1 and F-9 respectively without any significant difference ($p < 0.05$). According to Oluwatoyin *et al,* the swelling index of the microspheres also increased with increase in the concentration of starch in the polymer blend. The ibuprofen microspheres showed significant ($p < 0.01$) higher swelling index than those containing sodium alginate alone as the polymer.[19] The drug content was found to be between 28.00 % ± 0.00 to 43.00 % ± 2.83 for F-3 and F-9 respectively. This depicts that the formulation F-9 containing HPMC polymer gave the highest drug content, while F-3 gave the least drug content of 28.00 % ± 9.55.

Table 5: Swelling index and Drug content of ibuprofen microspheres (mean ± SD).

Formulation code	Swelling index (%)	Drug content (%)
F-1	65.50 ± 0.35	27.00 ± 0.71
F-2	71.50 ± 1.77	28.00 ± 1.41
F-3	72.50 ± 0.35	28.00 ± 0.00
F-4	73.00 ± 1.41	31.00 ± 2.83
F-5	78.00 ± 1.41	28.50 ± 1.77
F-6	79.00 ± 0.71	29.00 ± 0.71
F-7	78.50 ± 1.06	34.00 ± 1.41
F-8	81.50 ± 1.06	30.50 ± 0.35
F-9	85.00 ± 0.71	43.00 ± 2.83

3.8 *In vitro* release study

The drug release profile of the ibuprofen floating microspheres are found in Table 6. The drug release profile of the floating microspheres increased with time. According to Huang *et al*, for controlled release preparations, an initial high rate of drug release is usually observed at the beginning of the controlled release process which could be due to a number of mechanisms such as surface desorption, pore diffusion and lack of a diffusion barrier to regulate the diffusion process.[20] The results obtained showed that the time taken for at least 20 % of the ibuprofen release were higher than 60 minutes. This indicated that the ibuprofen microspheres did not show any sign of burst release, thus indicating that they might had been embedded in the microspheres.[21] From the results obtained, formulation F-1 gave the highest cumulative drug release of 70.54 % ± 2.47 at 240 minutes, while formulation F-5 containing HPMC provided the lowest release of 28.80 % ± 0.85 over the same period of time.

Table 6: *In vitro* drug release profile of ibuprofen microspheres (mean ± SD)

Time (mins)/Batches	F-F-1	F-2	F-3	F-4	F-5	F-6	F-7	F-8	F-9
0	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00
15	2.15 ± 0.11	1.80 ± 0.14	2.35 ± 0.11	2.45 ± 0.11	2.20 ± 0.28	2.25 ± 0.18	2.40 ± 0.07	2.75 ± 0.53	2.15 ± 0.25
30	8.15 ± 0.18	8.35 ± 0.25	7.80 ± 0.42	7.50 ± 0.28	7.75 ± 0.18	8.15 ± 0.25	7.95 ± 0.04	7.75 ± 0.46	8.20 ± 0.14
60	17.50 ± 1.77	7.55 ± 3.15	17.85 ± 2.02	14.85 ± 0.11	7.70 ± 3.46	7.40 ± 3.11	16.45 ± 1.03	17.15 ± 1.52	11.75 ± 0.04
90	35.50 ± 0.35	15.35 ± 0.25	35.50 ± 0.35	26.75 ± 1.94	14.80 ± 0.28	15.75 ± 0.74	36.65 ± 0.95	26.25 ± 1.59	15.20 ± 0.35
120	47.00 ± 0.00	13.45 ± 3.22	42.00 ± 0.71	36.75 ± 0.53	17.80 ± 0.14	17.60 ± 0.28	36.10 ± 0.07	34.10 ± 1.34	17.45 ± 0.39
150	52.00 ± 3.54	21.30 ± 0.92	49.50 ± 1.77	40.00 ± 0.00	19.75 ± 0.18	19.75 ± 0.04	43.00 ± 2.12	42.30 ± 2.62	19.70 ± 0.07
180	55.00 ± 3.54	23.55 ± 0.39	54.50 ± 1.77	49.55 ± 1.80	13.10 ± 0.79	23.80 ± 0.14	51.25 ± 0.88	51.00 ± 0.71	25.00 ± 1.41
210	61.50 ± 1.06	25.60 ± 0.28	64.00 ± 2.83	53.00 ± 2.12	25.35 ± 0.04	25.90 ± 0.42	57.40 ± 0.42	54.50 ± 2.47	32.80 ± 1.56
240	70.50 ± 2.47	30.55 ± 0.25	61.00 ± 0.00	61.90 ± 1.34	28.80 ± 0.85	28.80 ± 2.62	61.20 ± 0.57	64.50 ± 3.18	39.55 ± 0.67

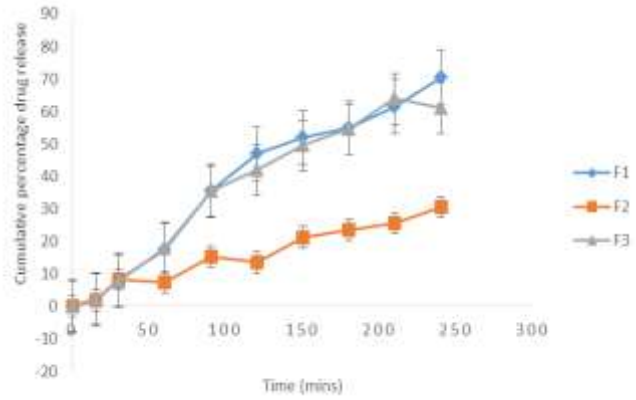

Fig. 1: Cumulative percentage drug release of batches F-1 to F-3.

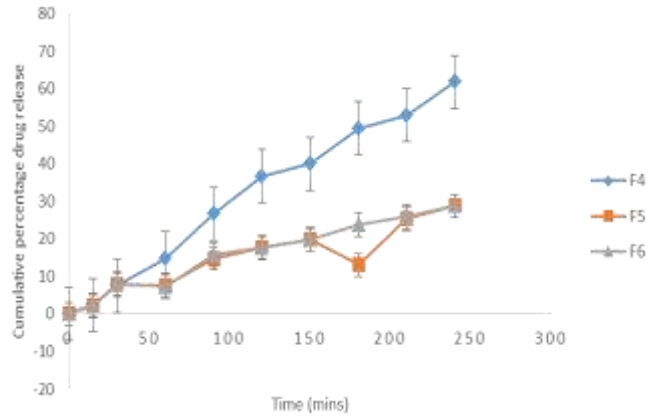

Fig. 2: Cumulative percentage drug release of batches F-4 to F-6.

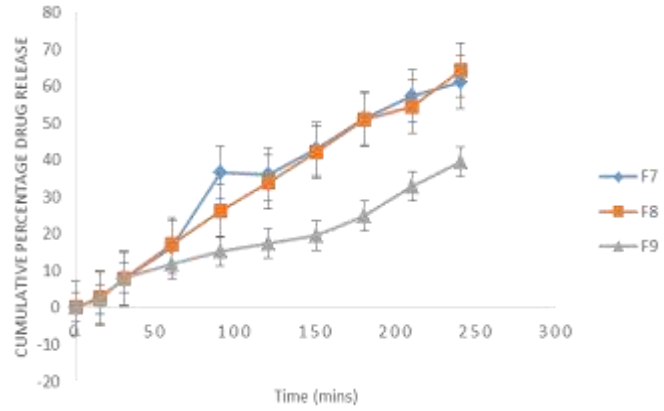

Fig. 3: Cumulative percentage drug release of batches F-7 to F-9.

3.9 Scanning electron microscopy

The ibuprofen microspheres were discrete and spherical in shape. Formulations F-1 and F-2 showed smoother surfaces than the microspheres that were formulated with HPMC (F-6 and F-7). According to Letful *et al,* the presence of guar gum on the surface of the microspheres might cause a slightly rough surface as guar gum might interfere with the cross-linking of alginate by calcium ions.[22]

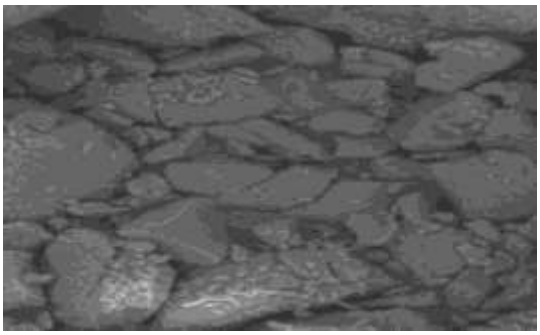

Fig. 4: SEM of batch F-1.

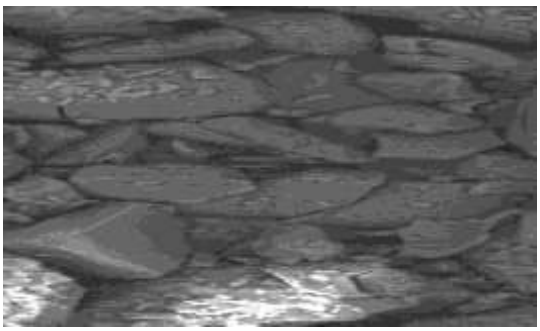

Fig. 5: SEM of batch F-2.

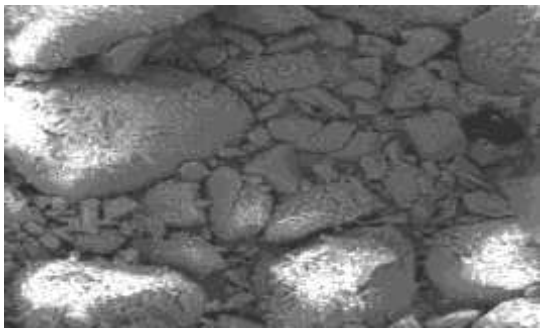

Fig. 6: SEM of batch F-7.

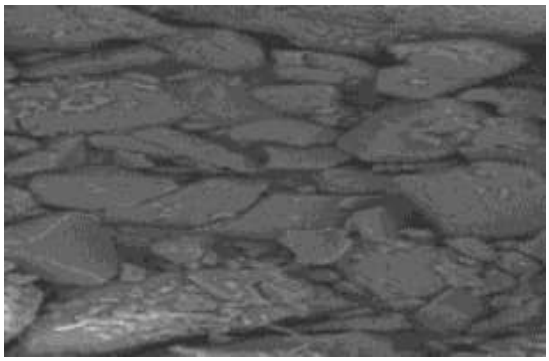

Fig. 7: SEM of batch F-8.

3.10 FTIR Spectroscopy (drug- excipient compatibility studies)

Fig.8 showed the characteristic peaks of ibuprofen at 3802.09, 2606.8, 2457.3, 1865.8 and 1421.0 cm $^{-1}$ corresponding to O-H single bond stretch, C-H single

bond stretch, nitriles and carbenes triple bond, C=O, C=C, C=N double bond and C-O, C-C single bond respectively.

Fig. 9 shows the characteristics peaks of ethyl cellulose at 3925.2, 3245.2, 2582.1, 1997.8 and 1468.4 cm^{-1} corresponding to O-H, N-H single bond stretch, C-H single bond stretch, carbenes triple bond, C=O, C=C double bond and C-C, C-O single bond respectively. According to Sunil *et al,* the spectrum of EC showed characteristic peaks at 3390 and a band at 1636 cm^{-1} corresponding to the stretching and bending modes of the surface hydroxyls. The peak at 2905 cm^{-1} belongs to the asymmetrically stretching vibration of C-H in a pyramid ring and the broad absorption peak at 1059 cm^{-1} is attributed to the C-O of cellulose.[23]

Fig. 10 shows the characteristic peaks of sodium carboxy-methyl cellulose at 3852.3, 3169.1, 2554.2, 1993.8 and 1495.6 cm^{-1} corresponding to O-H, N-H single bond stretch, C-H single bond stretch, nitriles and carbenes triple bond, C=O, C=C double bond and C-O, C-C single bond respectively. According to Mastiholimath *et al*, the spectrum of Na-CMC showed characteristic peaks at 3700 cm^{-1} indicating the presence of –OH stretching bond. The strong bonds at 1093, 459 and 798 cm^{-1} were associated to the asymmetric and symmetric Si-O—Si stretching vibration bonding.[24]

Fig.11 shows the characteristic peaks of HPMC at 3675.0, 3269.4, 2434.3, 1900.5 and 1428.8 corresponding to –O-H, single bond stretch, C-H single bond stretch, nitriles and carbenes triple bond, C=O C=C double bond, C-O, C-C single bond respectively. According to Shoufeng *et al;* the spectrum of starch showed characteristics peaks at 3448 for –OH stretching, 2930 for –CH stretching, 1646 for C-O bending associated with OH group, and 1381 cm^{-1} associated with –CH symmetric bending.[25]

Fig. 12 shows the characteristic peaks of calcium chloride at 3900.1, 3143.6, 2427.5, 1873.9 and 1454.0 corresponding to –OH, -NH single bond stretch, -CH single bond stretch, C=O, C=C and C-O, C-C single bond respectively. According to Yogesh *et al*, the twin peaks at 1577 and 1466 cm^{-1} were attributed to asymmetric carbohydrate (-COO) stretching vibration and symmetric carbohydrate vibration respectively, while peaks at 2917 and 2850 cm^{-1} were attributed to the –CH stretching vibration.[26]

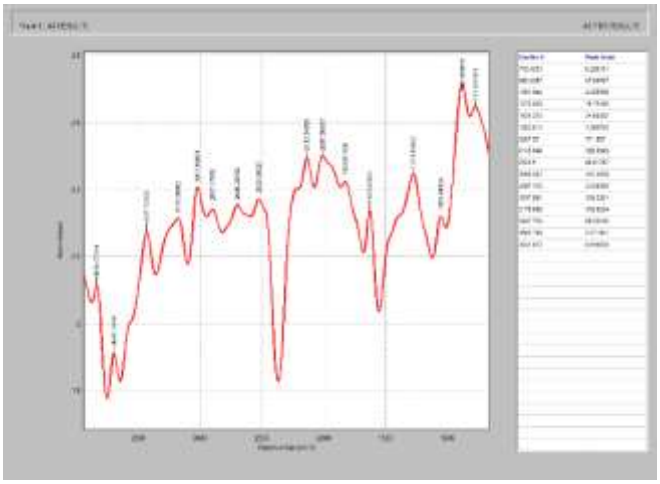

Fig. 8: FTIR spectrum of Metronidazole.

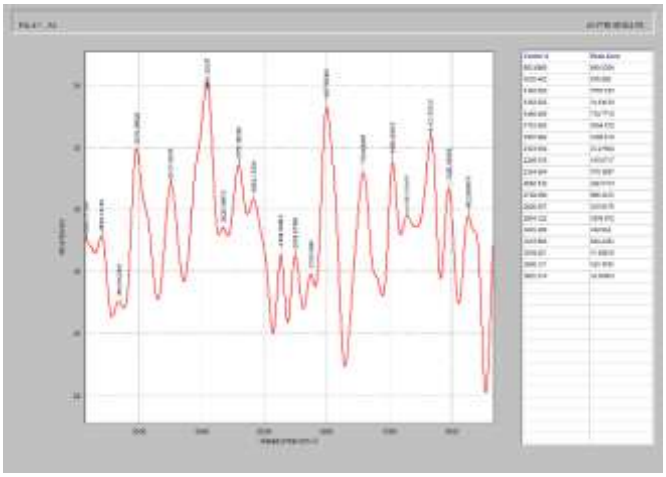

Fig. 9: FTIR spectrum of ethyl cellulose.

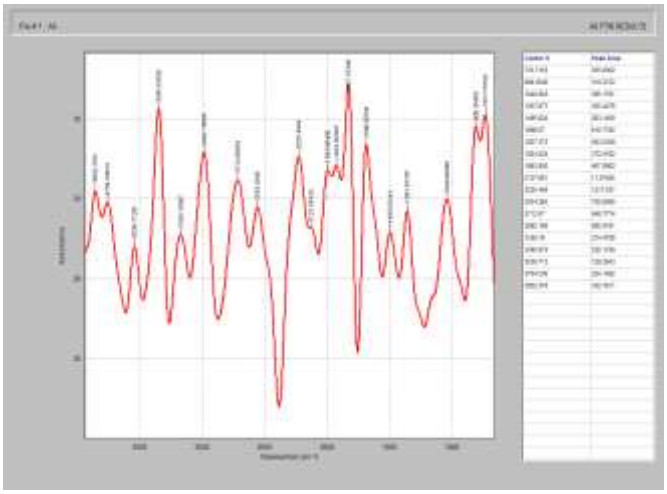

Fig. 10. FTIR spectrum of Na-CMC

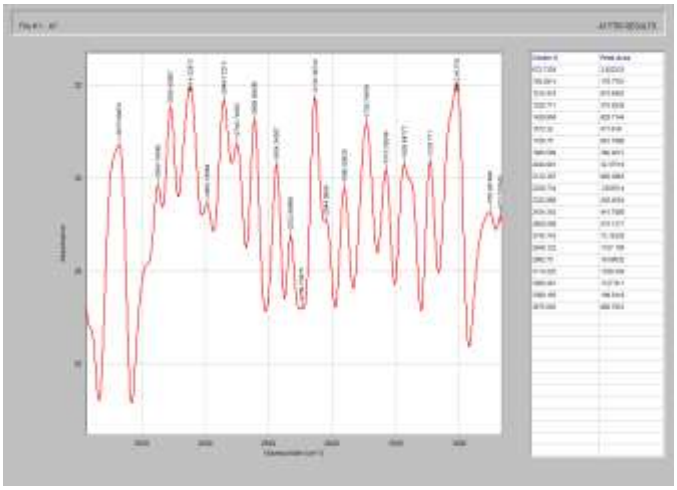

Fig. 11: FTIR spectrum of HPMC.

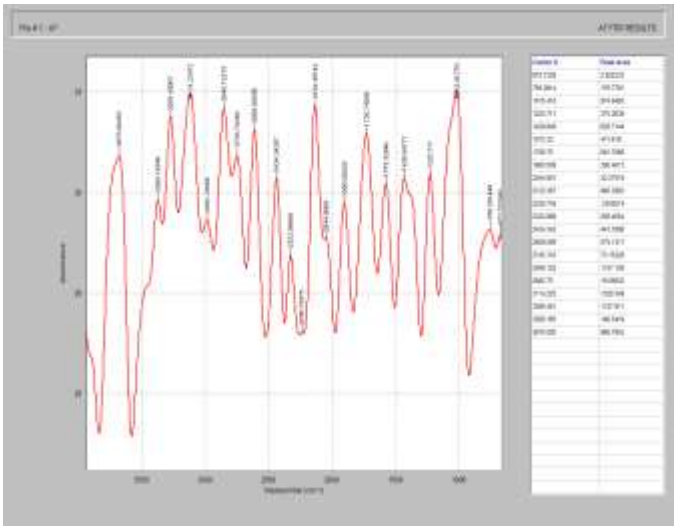

Fig. 12: FTIR spectrum of Calcium chloride.

CONCLUSION

The idea of formulating floating microspheres containing ibuprofen gave a suitable practical approach that achieved a prolonged therapeutic effect by releasing the active drug over an extended period of time.

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CONFLICT OF INTEREST

Authors declare no conflict of interest

ABBREVIATIONS

GIT: Gastrointestinal Tract
FDDS: Floating Drug Delivery Systems
Na-CMC: Sodium Carboxylmethyl Cellulose
FTIR: Fourier Transform Infrared
EC: Ethyl Cellulose
HPMC: Hydroxypropyl Methyl Cellulose
C.I: Carr's Index
H.R: Hausner's Ration

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